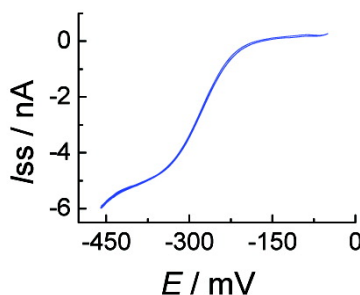
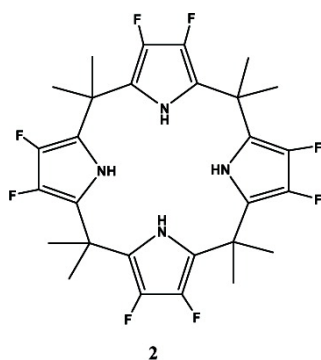


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Anion Transfer at a Micro-Water/1,2-Dichloroethane Interface Facilitated by β -Octafluoro-*meso*-octamethylcalix[4]pyrrole

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Due to its potential importance in drug delivery and value in understanding ion transport across biomembranes, facilitated ion transfer (FIT) at liquid/liquid (L/L) interfaces has been widely investigated since the pioneering work of Koryta in 1979.¹ To date, most reports in this area have focused on facilitated cation transfer (FCT).² In spite of great advances in the area of anion recognition chemistry,³ few studies involving facilitated anion transfer (FAT) at L/L interfaces have been published,⁴ and even fewer reports regarding a kinetic analysis of the underlying phenomenon have appeared.^{4j} Given this lack of available dynamic information, we have used electrochemical methods to study the thermodynamic and kinetic transfer behavior of several monovalent anions at a micropipet-supported micro-water/1,2-dichloroethane (μ -W/DCE) interface. Little effective transfer is observed in the absence of an additive or when calix[4]pyrrole **1** is used as a potential transfer agent (cf. Figure 1A for structures). On the other hand, in the presence of β -octafluoro-*meso*-octamethylcalix[4]pyrrole **2** effective transfer of Cl⁻, Br⁻, NO₂⁻, and CH₃CO₂⁻ (Ac⁻) at the μ -W/DCE interface is observed. However, the kinetic rate constants are 1–2 orders of magnitude smaller than those for analogous FCT processes involving crown ethers as the ion receptor.²

In simplified terms, FAT at a μ -W/DCE interface involves the following process:



where A⁻ and R represent a monovalent anion and the putative facilitating anion transfer agent (e.g., receptors **1** or **2**, respectively). Key goals were to (i) find a receptor that would allow for facilitated anion transfer and (ii) determine the transfer kinetics of anion transport in the event that such a transfer agent is identified.

Anions are relatively challenging substrates for FIT studies. They are usually larger than isoelectronic cations.^{3,5,6} They thus have lower charge densities and are less prone to bind strongly to potential transfer agents through electrostatic interactions. Compared to cations, inorganic anions also have larger Gibbs energies of transfer.⁴ Therefore, transfer of most free inorganic anions occurs at the negative limit or outside of the potential window accessible using typical μ -W/DCE setups. This means that it is inherently harder to find a receptor for FAT than it is for FCT.² To date, only a limited number of systems have been used to effect FAT and only when relatively large L/L interfaces (on the order of cm²) were employed.⁴ We sought to explore whether FAT could be achieved in a μ -W/DCE setup using a calix[4]pyrrole as the receptor.⁷

Calix[4]pyrroles, such as **1**, have attracted much attention as anion receptors over the past decade.⁵ In organic solution and solid phase, calix[4]pyrroles can bind anions such as F⁻, Cl⁻, Br⁻, I⁻, and H₂PO₄⁻ effectively.⁵ β -Octafluorocalix[4]pyrrole **2** was shown

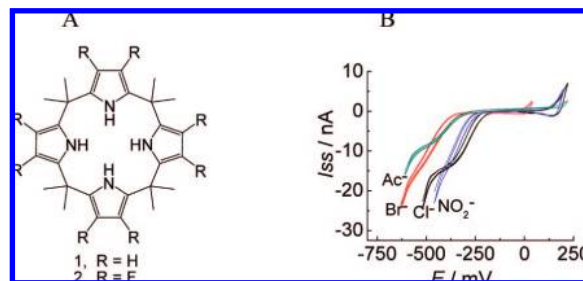


Figure 1. (A) Structures of calix[4]pyrroles **1** and **2**. (B) Cyclic voltammograms for transfer of Cl⁻, Br⁻, NO₂⁻, and Ac⁻ at a μ -W/DCE interface facilitated by **2** using Cell 1 ($x = 2$, $y = 100$) and a pipet radius, r , of 20, 20, 22, and 10 μ m, respectively. The scan rate is 50 mV/s.

Table 1. Thermodynamic Data for Anion Transfer Facilitated by **2**

anion	$\Delta_s^\circ \phi_{AR}^\circ$ (mV)	m/n	$\log \beta_{AR}^\circ$	D_R (cm ² /s)
Cl ⁻	-514 ^{11a}	1:1	8.26	4.2×10^{-6}
Br ⁻	-405 ^{11a}	1:1	6.52	3.6×10^{-6}
NO ₂ ⁻	-332 ^{11b}	1:1	3.28	3.2×10^{-6}
CH ₃ CO ₂ ⁻	-560 ^{11c}	1:1	5.77	3.0×10^{-6}
mean				$(3.5 \pm 0.5) \times 10^{-6}$

to have an increased anion binding affinity relative to the parent system **1** due to the presence of electron-withdrawing fluorine substituents.⁸ Calix[4]pyrroles have been studied in a number of applications, such as response elements in ion-selective electrodes.⁹ In this work, the FAT of small anions at the μ -W/DCE interface was studied electrochemically using both **1** and **2** as the possible receptors.

The μ -W/DCE electrochemical cell (Cell 1) employed for the cyclic voltammetry (CV) and differential pulse voltammetry (DPV) measurements (Supporting Information (SI)) is as follows:



Of the anions tested, four (Cl⁻, Br⁻, NO₂⁻, and Ac⁻) displayed reasonable steady-state voltammograms within the potential window when **2** was used as the receptor (Figure 1B). The negative current reflects the transfer of anions from the aqueous to the DCE phase.¹⁰ In the case of F⁻, which has the strongest interaction with **2** in organic media,⁵ no wave corresponding to its facilitated transfer can be observed within the potential window. This finding likely reflects that F⁻ is very hydrophilic and retained in the aqueous phase. From the CV and DPV curves, the thermodynamic parameters (such as the diffusion coefficient D_R , the stoichiometric ratio m/n , and the association constant $\log \beta_{AR}^\circ$) can be easily obtained (SI); these values are listed in Table 1.

Bard et al. have developed a three-point method to determine the kinetic parameters of a heterogeneous electron transfer reaction

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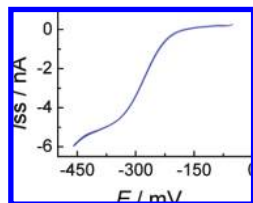


Figure 2. Steady-state voltammogram of Cl^- transfer at a $\mu\text{-W/DCE}$ interface facilitated by **2** using Cell 1, where $x = 2$, $y = 5000$. The scan rate is 50 mV/s, and $r = 7 \mu\text{m}$.

from a quasi-steady state voltammogram.¹² Three parameters, namely, the half-wave potential $E_{1/2}$ and the quartile potentials $E_{1/4}$ and $E_{3/4}$, which can be obtained experimentally, are used to determine the transfer coefficient (α) and the standard rate constant (k°).¹³ However, it is subject to several caveats, including the need for well-defined steady state voltammograms. In the particular case of the reaction shown in eq 1, if the three potentials obtained at a W/DCE interface cannot satisfy the conditions that (i) $|\Delta E_{1/4}| = E_{1/2} - E_{1/4} \geq 30.5 \text{ mV}$, (ii) $|\Delta E_{3/4}| = E_{3/4} - E_{1/2} \geq 31.0 \text{ mV}$, and (iii) $|\Delta E_{3/4}| \geq |\Delta E_{1/4}|$, then the reaction at the $\mu\text{-W/DCE}$ interface is reversible and no kinetic data can be obtained using this method. Reversibility depends on k° and the mass transport rate (k_d); therefore, it is useful to increase the k_d , e.g., by using a smaller nano-L/L interface, to determine even larger k° values. This is what has been done for the previously reported FCT studies.¹³ In accord with what would be expected based on eq S1 (SI), the $E_{1/2}$ was found to shift to a more positive potential as the anion concentration increased. Using this approach, well-defined steady state voltammograms for Cl^- and Ac^- could be observed (cf. Figure 2 for Cl^-). This allowed k° values of 2.11 ± 0.90 and 0.75 ± 0.50 ($\times 10^{-2} \text{ cm/s}$), as well as α values of 0.57 ± 0.07 and 0.62 ± 0.04 for Cl^- and Ac^- , respectively, to be determined (see Table T1 in the SI). Efforts to extend this analysis to Br^- and NO_2^- failed since their waves proved too close to the negative end of potential window.

From these results it is clear that the FAT at the W/DCE interface mediated by **2** is much slower than analogous FCT processes observed for alkali metal cations using crown ether as receptors.¹³ While too many differences exist to allow for direct comparisons, there are a number of likely explanations for this apparent dichotomy. First, most anion binding agents, including **2**, generally display lower affinities for their targeted anions than do receptors for similarly sized cations. While an obvious oversimplification, such a thermodynamic disadvantage is likely to translate into slower ion transfer kinetics.¹⁴ Second, anions of similar charge and size are usually characterized by higher hydration energies (ΔG_{hyd}); cf. e.g., $\text{F}^- -465/\text{Na}^+ -365 \text{ kJ mol}^{-1}$, $\text{Cl}^- -340/\text{K}^+ -295 \text{ kJ mol}^{-1}$.¹⁵ According to the Marcus theory,¹⁶ an ion transfer reaction at an L/L interface involves initial desolvation of an ion from the first phase and then concerted solvation by the second phase. The higher the hydration energy, the harder it is to overcome this barrier. This thermodynamic “penalty” (in the case of anions) is likely to be reflected in slower facilitated ion transfer kinetics. Anions also usually display shorter solvation times than cations (on the order of picoseconds and nanoseconds, respectively);¹⁷ however, the effect of this difference on the FIT kinetics is unclear. Nevertheless, this clear difference between cation and anion behavior is of inherent interest and could be useful in the design of ion sensors.

In summary, we have demonstrated the facilitated ion transfer of four anions, namely Cl^- , Br^- , NO_2^- , and Ac^- , by receptor **2** at a $\mu\text{-W/DCE}$ interface. We have also shown for the first time that the dynamics of this process can be studied by micropipet voltammetry. Studies such as these are expected to be useful in understanding the mechanism of anion transport at soft interfaces and for the design of yet-improved anion receptors and carriers.

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Supporting Information Available: Details of experiments and electrochemical measurements. This material is available free of charge via the Internet at <http://pubs.acs.org>.

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